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Atoms First | OpenStax balanced:

$2\text{Ag}(s) + \text{Cu}^{2+}(aq) \rightarrow 2\text{Ag}^{+}(aq) + \text{Cu}(s)$
 $E^{\circ} = E^{\circ}_{\text{red}}(\text{Cu}^{2+}) - E^{\circ}_{\text{ox}}(\text{Ag})$
 $= 0.34\text{V} - 0.80\text{V} = -0.46\text{V}$. Not

spontaneous. ox: $[\text{Ni}(s) \rightarrow \text{Ni}^{2+}(aq) + 2e^{-}]$
 $\times 5$; red: $[5e^{-} + 8\text{H}^{+}(aq) + \text{MnO}_4^{-}(aq) \rightarrow 4\text{H}_2\text{O}(l) + \text{Mn}^{2+}(aq)]$

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+ + 4H₂O] × 2. balanced: 5Ni(s) +

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balanced: 2Ag(s) + Cu²⁺ + (aq) → 2Ag +
(aq) + Cu(s) E^o = E^o ox(Ag) + E^o
red(Cu²⁺) = -0.80V + 0.34V = -
0.46V. Not spontaneous. ox: [Ni(s) →

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$\text{Ni}^{2+} + 2\text{e}^{-}] \times 5$; red: $[5\text{e}^{-} + 8\text{H}^{+} + \text{MnO}_4^{-} \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}] \times 2$.

balanced: $5\text{Ni(s)} + 16\text{H}^{+}(\text{aq}) + 2\text{MnO}_4^{-}(\text{aq}) \rightarrow 5\text{Ni}^{2+}(\text{aq}) + 2\text{Mn}^{2+}(\text{aq}) + 8\text{H}_2\text{O}$.

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Key Ideas Answers 14. Because protons and neutrons reside in the nucleus of atoms, they are called nucleons. 16. There are two forces among the particles within the nucleus. The first, called the electrostatic force, is the force between electrically charged particles. The second force,

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a) the energy given off when reactants collide b) *an energy barrier between reactants and products the heat released in a reaction* c) generally very high for a reaction that takes place rapidly

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CH₃ - CH₂ - NH₂ is classified as a. A) primary amine. B) secondary amine.

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