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~~Chapter 17 Calorimetry: Crash Course Chemistry #19 Gibbs Free Energy - Equilibrium Constant, Enthalpy \u0026 Entropy - Equations \u0026 Practice Problems Chemical Kinetics Rate Laws - Chemistry Review - Order of Reaction \u0026 Equations Enthalpy Change of Reaction \u0026 Formation - Thermochemistry \u0026 Calorimetry Practice Problems~~
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You must show all work and setup
for this to count as extra credit on
your test (+3 points) 15.1 - 15.2

Heat, Calorimetry, and Enthalpy.

Make the following conversions:

444 cal to joules. 1.8 kJ to joules.

0.45 kJ to calories.

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THERMOCHEMISTRY SECTION

17.1 THE FLOW OF ENERGY-HEAT

AND WORK (pages 505-510) This

section explains the relationship

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distinguishes between heat capacity and specific heat. ~ Energy Transformations (page 505) 1. What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions?
thermochemistry

~~THERMOCHEMISTRY~~

Chapter 17 Thermochemistry 183

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK

(pages 505–510) This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat. Energy Transformations (page 505) 1. What area of study in chemistry is concerned with the heat transfers

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~~SECTION 17.1 THE FLOW OF
ENERGY HEAT AND WORK (pages
505-510)~~

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Guide Thermochemistry $Q = m C$
 ΔT $\Delta H = -q$ $C = 4184 \text{ J moles } ^\circ\text{C}$
1 When 150-g sample of KCl
dissolves in 650 g of water in a
calorimeter, the temperature
drops from 310°C to 205°C
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study of energy changes that
occur during chemical reactions
and changes in state. ...

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thermochemistry this chapter
explores ideas related to heats of
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CHAPTER 17: FREE ENERGY AND
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NIVALDO J. TRO For this chapter
we will review some of Chapter 6,
Thermochemistry, and study a
part of Chapter 6 which we
omitted in General Chemistry I.
Everything in the universe is
energy or matter.

~~Chapter 17 Review Master
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