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the outermost and highest energy electrons in an atom; an atom has eight at most. pauli exclusion principle. no two electrons may have the exact same energy state, two electrons may occupy one orbital but they must have opposite spins; cannot draw two arrows with the same direction in an orbital. Spin.

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fixed. energy. levels. Electrons are never between energy levels or energy shells. An electron must have . just the right amount of energy. to jump from one level to another. A Chapter 13 Electrons in Atoms Last modified by:

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Chapter 18 Study Guide 1. Why do atoms gain or lose electrons? To get a stable electron ____ configuration

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____.(full outer shell - octet)) 2. Which group on the periodic table contains elements with stable electron configurations? ____noble ____ ____gases ____ Group 18 3. If you have a full outer energy level you have a stable __octet __ or __configuration __. 4.

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Organization of Electrons in Atoms – Introductory ...

You might have been taught that atoms are the smallest particles on Earth, but inside of atoms are even smaller

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things called electrons. ... The number of electrons in an atom The way atoms bond ...

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